

Calculating Average Rates of Reaction

You are required to show work for all calculations.

1. During the combustion of methane, CH_4 , shown by the reaction



The concentration of methane was measured at various time intervals and the following results were obtained:

Time (s)	$[\text{CH}_4]$ ($\text{mol} \cdot \text{l}^{-1}$)
10	2.40
20	1.20
30	0.80
40	0.60

Calculate the average rate of loss of methane during the 10 to 40 second time period.

Solution:

$$\begin{aligned}
 \text{rate} &= \frac{\Delta[\text{CH}_4]}{\Delta t} \\
 &= \frac{[0.60 - 2.4]\text{mol}}{(40 - 10)\text{s} \times \text{L}} \\
 &= \frac{-1.8\text{mol}}{30\text{s} \times \text{L}} \\
 &= -0.06 \frac{\text{mol}}{\text{L} \times \text{s}}
 \end{aligned}$$

Note: the answer is negative because we are calculating the **rate of loss of methane**.

Calculating Reaction Rates using Stoichiometry

Example:

In the following decomposition reaction,



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Barry McBride

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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