

Stoichiometry Worksheet #6: Mixed Practice

OHS Chemistry 4/14/2006

Check To Make Sure Given Equations Are Balanced!

1. When aluminum is heated in oxygen, aluminum oxide is formed:



- a) Determine the limiting reactant if 25.0g of aluminum metal and 20.0g of oxygen are reacted according to the above equation.
b) What mass of the oxide would be obtained from same reaction?
2. 50.0g of oxygen are available for synthesis with 25.0g of carbon.



- a) Which reactant is limiting?
b) How much carbon dioxide would be produced in this reaction?
3. a) How many grams of carbon dioxide can be obtained from the reaction of 100.0g of sulfuric acid and 100.0g of calcium carbonate?
b) Which reactant is in excess and how much of it would remain?



4. In the combustion of hydrogen sulfide with oxygen, sulfur dioxide and water are produced.

Will 64.29 g be enough oxygen to completely burn 35.0g of hydrogen sulfide?
(Hint: consider the limiting reactant)

5. When steam (gaseous water) is passed over iron, hydrogen gas and iron (III) oxide are formed. What mass of steam would be needed to react completely with 100.0g of iron? (Hint: what needs to be done to the equation first?)



6. In a double replacement reaction where 100.0g ammonium hydroxide reacts with 75.0g of copper (II) nitrate, what mass of ammonium nitrate would be generated?

7. When ammonia (NH₃) is burned in oxygen, free nitrogen gas and water are produced

- a) What mass of ammonia will react completely with 35.7 g of oxygen?
b) What mass of nitrogen gas is formed?



8. When sodium carbonate reacts with hydrochloric acid, the carbonic acid that is formed immediately breaks down into carbon dioxide and water.(see reaction below) What mass of sodium chloride must have been originally present if 9.82 g of carbon dioxide were obtained in this way?



9. How much copper metal can be obtained by the single replacement reaction between 30.0 g of copper (I) nitrate and 30.0g of iron metal?

10. What mass of sulfuric acid will be needed to react completely with 34.2 g of ammonia in the production of ammonium sulfate?

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