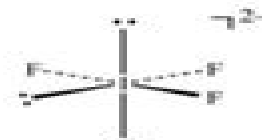


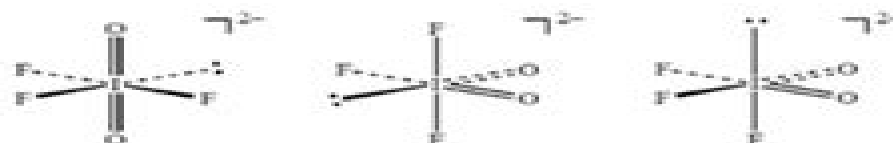
$(\text{CH}_3)_3\text{NAl}(\text{CH}_3)_3$ relative to $(\text{CH}_3)_3\text{PAl}(\text{CH}_3)_3$ (1.978 pm vs. 1.973 pm), these lengths are not statistically different when their standard deviations are considered.

Data for $(\text{CH}_3)_3\text{NAl}(\text{CH}_3)_3$ from T. Gelbrich, J. Sieler, U. Dümichen, *Z. Kristallogr.*, **2000**, 215, 127. Data for $(\text{CH}_3)_3\text{PAl}(\text{CH}_3)_3$ from A. Almenningen, L. Fernholt, A. Haaland, J. Weidlein, *J. Organomet. Chem.*, **1978**, 145, 109.

- 3.20 IF_3^{2-} has three lone pairs and three bonds. Overall, this ion is predicted to be T-shaped, with bond angles slightly less than 90° .



- 3.21 a. There are three possibilities:



- b. The third structure, with the lone pair and double bonds in a facial arrangement, is least likely because it would have the greatest degree of electron-electron repulsions involving these regions of high electron concentrations.

The second structure, which has fewer 90° lone pair–double bond repulsions than the first structure, is expected to be the most likely. Experimental data are most consistent with this structure.

- c. One possibility: XeO_2F_3^-

- 3.22 a. Three unique arrangements of the nonbonding pairs and the oxygen atom are possible in XeOF_3^- when the octahedral electron-group geometry is considered: (1) *trans* nonbonding pairs; (2) one pair *trans* and one *cis* the oxygen atom; and (3) both pairs *cis* the oxygen atom.



A square-planar structure (1) is expected to minimize *lp-lp* repulsions relative to structures 2 and 3 with unfavorable 90° *lp-lp* interactions. A low temperature Raman spectroscopic study coupled with quantum-chemical calculations of shock-sensitive salts of XeOF_3^- confirms this prediction (D. S. Brock, H. P. A. Mercier, G. J. Schrobilgen, *J. Am. Chem. Soc.*, **2010**, 132, 10935).

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