

Chemistry Preliminary Exam Notes

Module 1 - Properties And Structure Of Matter

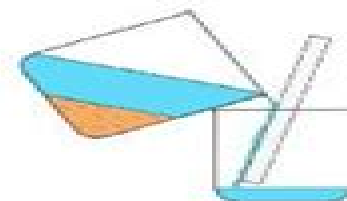
How Do The Properties Of Substances Help Us To Classify And Separate Them?

WHATS IN THIS LESSON (PART 1 + PART 2):

- Explore homogeneous mixtures and heterogeneous mixtures through practical investigations:
 - Using separation techniques based off physical properties.
 - Calculating percentage composition by weight of component elements and/or compounds.

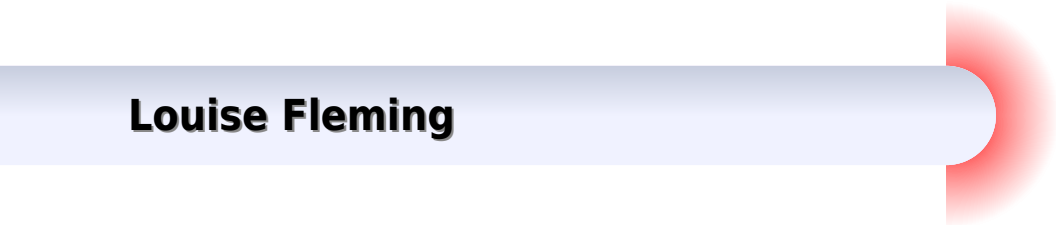
Separation Of Mixtures (Part 1):

- 2 Types Of Matter:
 - Pure Substance
 - Mixture
- Pure Substance
 - 2 Types
 - Elements
 - Compounds
- Mixtures
 - 2 Types
 - Homogeneous
 - Heterogeneous
 - Variable in composition
 - Can be separated by physical means.
 - E.G. Based on their size, base on temperatures they boil at.
- **Homogeneous:** A mixture with indistinguishable components.
 - E.G. Tea with sugar, air, tap water.
- **Heterogeneous:** A mixture where the components are physically distinguishable.
 - E.G. Oil and water, wet sand, bread, gravel, etc.
- **Phase:** State of Matter
 - E.G. Solid, liquid, gas, etc.
- Homogeneous Mixtures
 - Have the same consistency throughout.
 - All components are in (or dissolved into) a single phase/state of matter.
- Heterogeneous Mixtures
 - Differing consistency throughout
 - All components don't have to be in the same phase/state of matter.
- **Immiscible:** Two liquids that will not dissolve into one another. They do not form distinct layers.
- Separation Techniques
 - Decantation
 - **Theory Summary:** With immiscible liquids or mixtures with suspensions they will separate out according to density (with the most dense at the bottom and form layers. The top or bottom layer can be poured off - This is called decantation.
 - **Explanation:**
 - Physical property that separation is based on: density.
 - The mixture is allowed to settle (this may take longer with solids suspended in liquids).
 - One layer is poured off very carefully.
 - There are limitations to the purity of the separated components **BUT** it is quick.



Preliminary Chemistry Revision

Louise Fleming



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