

Atomic structure and Nuclear Decay- Test B

Study Guide

- The disintegration of the radioisotope radium-226 produces an isotope of the element radon and alpha radiation. Write a balanced equation for this transformation
- The following radioisotopes are beta admitters. Write balanced nuclear equations for their decay process.
 - Carbon-14
 - strontium-90
 - potassium-40
 - nitrogen-13
- How are the mass number and atomic number of a nucleus affected by the following:
 - B- decay
 - α -decay
 - γ -emission
 - electron capture
 - positron emission
- Complete the following table.

3 nuclear particles/ Types of Radiation	Composition	Symbol	Mass in amu	Charge	Penetrating Power	Shielding
Alpha						
Beta						
Gamma			0			

- List three methods used to detect radiation.
- Define half-life
- Define transmutation and list two causes.
- List three uses for nuclear radiation
- What s the band of stability?
- Are all nuclei equally stable? What determines a nuclei's stability?
- What is mass defect? Why does it occur? What happens to the mass?
- List the six flavors of Quarks.
- What is a proton made of?
- What is a neutron made of?
- Which is heavier a neutron or a proton? Why?
- What are electrons made of?
- What is the difference between fission and fusion?
- What unit is used to measure molar mass?
- What is Avogadro's number? What is it based on?
- Determine the molar mass for the following:
 - Al
 - Mg
 - H_2O
 - CO_2
 - $BaCl_2$
 - $Ca(ClO_3)_2$
- Determine the mass of the following in grams.
 - 1.366 mol of NH_3
 - 238 mol of arsenic
 - 4.5 mol of NaF

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Albert A Gayle



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