

Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)

1. Physical Science Basics

- What tool(s) would you want to use to find the:
 - Mass of an object? triple beam balance Basic SI Unit? kg
 - Volume of an object? ruler or water displacement method Basic SI Unit? cm³ or ml
 - Length of an object? Ruler Basic SI Unit? meter
 - Density of an object? graduated cylinder Basic SI Unit? ml
- What is the formula for calculating density? mass/volume
 - What is the density of a cube that is 25 g and has a volume of 3 cm³? 8g/cm³
 - Compare the density of solid H₂O to liquid H₂O, and explain what observable effect this has. solid H₂O less dense than liquid H₂O

2. Chemistry Basics

- Element (def.) substance made up of one kind of atom
- Atom (def.) basic building block of matter
- Compound (def.) a molecule that contains atoms or molecules of two or more different elements
- Molecule (def.) a combination of two or more of the same or different atoms that are bonded together

3. Model of the Atomic Evolution of the Atomic Theory

Atoms are so small that many scientists throughout history have created models to describe them.

Why have atomic theories changed over time? Better technology came around.

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron	Plum Pudding Model
Niels Bohr	1913	Electrons move <u>fixed orbits</u> around the atom's nucleus in specific layers, or <u>electron shells</u> . Atoms absorb or give off energy when the electrons move from one shell (<u>energy level</u>) to another.	Bohr Model
Ernest Rutherford	1909	1909 /Based on idea that positive reflects positive/ atom mostly empty space; electrons orbit randomly around positively charged nucleus	Planetary Model

4. Periodic Table

- The current periodic table is arranged by increasing atomic number.
- Reading the Periodic Table
 - An element's properties can be predicted from its location in the periodic table.
 - The atomic number represents the number of protons.
 - The atomic mass represents the number of protons and neutrons.

Physical Science Semester Study Guide

Gerald E. Buck



Physical Science Semester Study Guide:

A Study Guide for Physics II Gerald E. Buck, 1966 *A Student's Guide Through the Great Physics Texts* Kerry Kuehn, 2014-09-15 This book provides a chronological introduction to the science of motion and rest based on the reading and analysis of significant portions of Galileo's Dialogues Concerning Two New Sciences Pascal's Treatise on the Equilibrium of Fluids and the Weight of the Mass of Air Newton's Mathematical Principles of Natural Philosophy and Einstein's Relativity Each chapter begins with a short introduction followed by a reading selection Carefully crafted study questions draw out key points in the text and focus the reader's attention on the author's methods analysis and conclusions Numerical and laboratory exercises at the end of each chapter test the reader's ability to understand and apply key concepts from the text Space Time and Motion is the second of four volumes in A Student's Guide through the Great Physics Texts This book grew out of a four semester undergraduate physics curriculum designed to encourage a critical and circumspect approach to natural science while at the same time preparing students for advanced coursework in physics This book is particularly suitable as a college level textbook for students of the natural sciences history or philosophy It also serves as a textbook for advanced high school students or as a thematically organized source book for scholars and motivated lay readers In studying the classic scientific texts included herein the reader will be drawn toward a lifetime of contemplation **Resources in Education**, 1994 *United States Armed Forces Institute Catalog* United States Armed Forces Institute, 1963 **Medical Laboratory Technician Certification Study Guide 2025-2026** Lucas Blade Hebert, Master Laboratory Science Ace Your Certification Ready to excel in your MLT or MLS certification exam This comprehensive review guide provides everything you need to succeed from fundamental concepts to advanced clinical applications 700 Practice Questions Test your knowledge across all laboratory disciplines with detailed explanations that reinforce learning 26 Complete Chapters Master Hematology Chemistry Microbiology Immunology Blood Banking and more with systematic easy to follow content Real Clinical Cases Apply your knowledge through authentic laboratory scenarios that mirror actual practice High Yield Review Focus your study time on the most frequently tested concepts with targeted review sections Test Taking Strategies Learn proven techniques to maximize your performance on examination day From basic laboratory principles to complex diagnostic challenges this guide builds the knowledge and confidence you need for certification success Perfect for MLT and MLS candidates laboratory science students and practicing professionals seeking continuing education Transform your preparation Achieve your certification goals **Research in Education**, 1974 Science Course Improvement Projects National Science Foundation (U.S.), 1962 **DOD Pam** United States. Office of Armed Forces Information and Education, *Promotion Fitness Examination study guide*, 2003 **Directory of Distance Learning Opportunities** Modoc Press, Inc., 2003-02-28 This book provides an overview of current K 12 courses and programs offered in the United States as correspondence study or via such electronic delivery systems as satellite cable or the Internet The Directory includes over 6 000 courses offered by 154

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