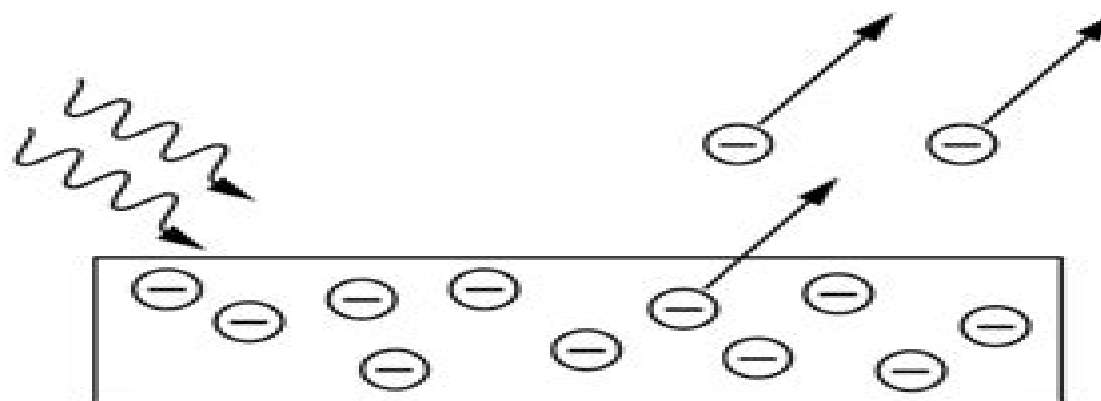


Photoelectric Effect Quiz



The diagram above represents the photoelectric effect for a metal. When the metal surface is exposed to light with increasing frequency and energy of photons, electrons first begin to be ejected from the metal when the energy of the photons is $3.3 \times 10^{-19} \text{ J}$.

1. Which of the following is closest to the frequency of the light with photon energy of $3.3 \times 10^{-19} \text{ J}$?

- (A) $5.0 \times 10^{-53} \text{ s}^{-1}$
 (B) $5.0 \times 10^{-16} \text{ s}^{-1}$
 (C) $5.0 \times 10^{14} \text{ s}^{-1}$
 (D) $5.0 \times 10^{52} \text{ s}^{-1}$

In photoelectric effect:

$$E = hf \quad \Rightarrow \quad f = \frac{E}{h} = \frac{3.3 \times 10^{-19} \text{ J}}{6.626 \times 10^{-34} \text{ J}\cdot\text{s}}$$

$$f = 4.98 \times 10^{14} \text{ s}^{-1} = 5.0 \times 10^{14} \text{ s}^{-1}$$

Frequency of the light is $5.0 \times 10^{14} \text{ Hz}$

$$E = 3.3 \times 10^{-19} \text{ J}$$

$$h = \text{Planck constant} = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$$

- 2.

Color	Wavelength
Red	647 – 760 nm
Orange	585 – 647 nm
Yellow	575 – 585 nm
Green	491 – 575 nm
Blue	424 – 491 nm
Violet	300 – 424 nm

Using the wavelength information provided above, what is the color of the light?

$$E = 3.3 \times 10^{-19} \text{ J}$$

$$\lambda = 400 \text{ nm}$$

$$h = 6.626 \times 10^{-34} \text{ J}\cdot\text{s}$$

$$c = 3 \times 10^8 \text{ m/s (velocity of light)}$$

$$E = hf = \frac{hc}{\lambda} \quad \Rightarrow \quad \lambda = \frac{hc}{E} = \frac{6.626 \times 10^{-34} \text{ J}\cdot\text{s} \times 3 \times 10^8 \text{ m/s}}{3.3 \times 10^{-19} \text{ J}}$$

$$\lambda = 6.0236 \times 10^{-7} \text{ m} = 6.0236 \times 10^2 \text{ nm} \left(\frac{10^9 \text{ nm}}{1 \text{ m}} \right) = 602.36 \text{ nm}$$

Color: Orange

Photoelectric Effect Quiz

**Barron's Educational Series, Robert
Jansen, Greg Young**



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