

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

**16****SOLUTIONS****SECTION 16.1 PROPERTIES OF SOLUTIONS (pages 471–477)**

*This section identifies the factors that affect the solubility of a substance and determine the rate at which a solute dissolves.*

**► Solution Formation (pages 471–472)**

Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2.

- Underline the condition that causes sugar to dissolve faster in water.
  - as a whole cube                      or            in granulated form?
  - when allowed to stand            or            when stirred?
  - at a higher temperature            or            a lower temperature?
- Name three factors that influence the rate at which a solute dissolves in a solvent.
  - agitation (stirring or shaking)
  - temperature
  - particle size
- Is the following sentence true or false? Finely ground particles dissolve more rapidly than larger particles because finer particles expose a greater surface area to the colliding solvent molecules. true

**► Solubility (pages 472–473)**

- Complete the following table showing the steps in a procedure to determine the total amount of sodium chloride that will dissolve in 100 g of water at 25°C.

| Procedure                                  | Amount Dissolved | Amount Not Dissolved |
|--------------------------------------------|------------------|----------------------|
| Add 36.0 g of sodium chloride to the water | 36.0 g           | 0.0 g                |
| Add an additional 1.0 g of sodium chloride | <u>0.2 g</u>     | <u>0.8 g</u>         |
| Determine the total amount that dissolves  | <u>36.2 g</u>    |                      |

- The amount of a substance that dissolves in a given quantity of solvent at a constant temperature is called the substance's solubility at that temperature.
- If a solution contains the maximum amount of solute for a given quantity of solvent at a constant temperature, it is called a(n) saturated solution.

# Properties Of Solutions Answers

**James M. Sangster**



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