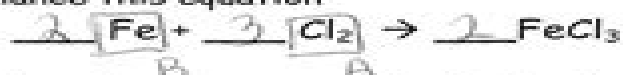


Stoichiometry Reflect #4

Learning Target... I can use stoichiometry to calculate amounts

1a. Balance this equation



1b. Determine how many moles of chlorine gas was used if the student weighed out 132 g of it.

$$132 \text{ g Cl}_2 \left(\frac{1 \text{ mol}}{70.90 \text{ g}} \right) = \boxed{1.86 \text{ moles Cl}_2}$$

1c. Determine the amount of moles of iron also used in this reaction.

$$1.86 \text{ moles A} \left(\frac{2 \text{ mol B}}{3 \text{ mol A}} \right) = \boxed{1.24 \text{ moles Fe}}$$

1d. Determine how many grams of iron were used in this reaction.

$$1.24 \text{ mol Fe} \left(\frac{55.85 \text{ g}}{1 \text{ mole}} \right) = \boxed{69.3 \text{ g Fe}}$$

2. Using the following chemical reaction:



and, knowing that 86 g of phosphorus trichloride was used in this experiment, how many molecules of water was also used?

$$86 \text{ g PCl}_3 \left(\frac{1 \text{ mole A}}{137.32 \text{ g}} \right) \left(\frac{3 \text{ mol B}}{1 \text{ mol A}} \right) \left(\frac{6.02 \times 10^{23} \text{ mc}}{1 \text{ mol B}} \right) = \boxed{1.1 \times 10^{24} \text{ mc H}_2\text{O}}$$

3. Challenge...

What is the percent yield of the following reaction if 4.0 g of oxygen gas was used in the lab, and 4.6 g of carbon dioxide was measured at the end of the experiment? (hint: you must calculate the theoretical yield first!!)



$$4.0 \text{ g O}_2 \left(\frac{1 \text{ mol A}}{32 \text{ g}} \right) \left(\frac{4 \text{ mol B}}{5 \text{ mol A}} \right) \left(\frac{44.01 \text{ g}}{1 \text{ mol}} \right) = 4.4 \text{ g CO}_2 \text{ theoretical}$$

$$\% \text{ yield} = \frac{4.6 \text{ g}}{4.4 \text{ g}} = \boxed{105\%}$$

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