MIXED GAS LAWS WORKSHEET - SOLUTIONS

 How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

n = <u>PV</u> = <u>(2.8 atm)(98 L)</u> = 11 moles of gas (0.0821 Latm/mol.K)(292 K)

If 5.0 moles of O₂ and 3.0 moles of N₂ are placed in a 30.0 L tank at a temperature of 25° C, what will the pressure of the resulting mixture of gases be? 25° C = 298 K

- O₂: P = <u>nRT</u> = (5.0 mol)(0.0821 Latm/molK)(298 K) = 4.1 atm V (30.0 L)
- N₂: $P = \frac{(3.0 \text{ mol})(0.0821 \text{ Latm/molK})(298 \text{ K})}{V} = 2.4 \text{ atm}$ V (30.0 L)

 $P_{Tot} = P_{02} + P_{N2} = 4.1$ atm + 2.4 atm = 6.5 atm Alternately you can substitute total moles of gas (8.0) as the total pressure depends on the total moles of gas - what type of gas is irrelevant

3) A balloon is filled with 35.0 L of helium in the morning when the temperature is 20.0°C. By noon the temperature has risen to 45.0°C. What is the new volume of the balloon?

 $T_1 = 20.0^{\circ} C = 293 \text{ K}, V_1 = 35.0 \text{ L}, T_2 = 45.0^{\circ} C = 318 \text{ K}, V_2 = 7$ $V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 \text{ L})(318 \text{ K})}{(293 \text{ K})} = 38.0 \text{ L}$

4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

n = <u>PV</u> = <u>(190 atm)(35 L)</u> = 260 moles of gas RT (0.0821 Latm/mol.K)(315 K)

5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in °C of the balloon?

T = <u>PV</u> = <u>(1 atm)(85 L)</u> = 296 K = 23°C nR (3.5 mol)(0.0821 Latm/mol.K)

6) CaCO₃ decomposes at 1200°C to form CO₂ gas and CaO. If 25 L of CO₂ are collected at 1200°C, what will the volume of this gas be after it cools to 25°C?

 $T_1 = 1200$ °C = 1473 K, $V_1 = 25$ L, $T_2 = 25$ °C = 298 K, $V_2 = 7$ $V_2 = V_1 T_2 = (25 \text{ L})(298 \text{ K}) = 5.1 \text{ L}$ (1473 K)

7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at 20.0°C is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is -20.0°C?

 $\begin{array}{ll} P_1 = 1.00 \text{ atm, } V_1 = 4.50 \text{ L, } T_1 = 20.0^{\circ}\text{C} = 293 \text{ K, } P_2 = 0.600 \text{ atm, } V_2 = ?, T_2 = -20.0^{\circ}\text{C} = 253 \text{ K} \\ V_2 = \underbrace{P_1 \, V_1 \, T_2}_{T_1 P_2} = \underbrace{(1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K})}_{(293 \text{ K})(0.600 \text{ atm})} = 6.48 \text{ L} \\ \end{array}$

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Holt Rinehart & Winston, Holt, Rinehart and Winston Staff

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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 The Electrical Review, 1888

Applied Mechanics Reviews, 1974 The Telegraphic Journal and Electrical Review, 1888 Digest; Review of Reviews Incorporating Literary Digest, 1898 Energy Research Abstracts, 1994 Journal of Research of the National Bureau of Standards, 1985 Journal of Research of the National Bureau of Standards United States.

National Bureau of Standards, 1985 Scientific and Technical Aerospace Reports, 1995 The Eclectic Magazine of Foreign Literature, Science, and Art, 1895

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