

Net Ionic Equations

Select the correct balanced **net ionic equation** for each of the following reactions, occurring in aqueous solution.

- ☒ $\text{K}_2\text{CO}_3 + \text{Ca}(\text{NO}_3)_2 \longrightarrow \text{CaCO}_3 + 2\text{KNO}_3$
- ☐ $2\text{NaBr} + (\text{NH}_4)_2\text{SO}_4 \longrightarrow 2\text{NH}_4\text{Br} + \text{Na}_2\text{SO}_4$
- ☐ $\text{CuSO}_4 + \text{CaCl}_2 \longrightarrow \text{CuCl}_2 + \text{CaSO}_4$

- A.** $\text{K}_2\text{CO}_3 + \text{Ca}^{2+} \longrightarrow 2\text{K}^+ + \text{CaCO}_3$
- B.** $\text{Ca}(\text{NO}_3)_2 + 2\text{K}^+ \longrightarrow \text{Ca}^{2+} + 2\text{KNO}_3$
- C.** $(\text{NH}_4)_2\text{SO}_4 + 2\text{Br}^- \longrightarrow 2\text{NH}_4\text{Br} + \text{SO}_4^{2-}$
- D.** $\text{SO}_4^{2-} + \text{Ca}^{2+} \longrightarrow \text{CaSO}_4$
- E.** $2\text{NH}_4^+ + 2\text{Br}^- \longrightarrow 2\text{NH}_4\text{Br}$
- F.** $2\text{NO}_3^- + 2\text{K}^+ \longrightarrow 2\text{KNO}_3$
- G.** $\text{SO}_4^{2-} + 2\text{Na}^+ \longrightarrow \text{Na}_2\text{SO}_4$
- H.** $2\text{Cl}^- + \text{Cu}^{2+} \longrightarrow \text{CuCl}_2$
- I.** $\text{CuSO}_4 + \text{Ca}^{2+} \longrightarrow \text{Cu}^{2+} + \text{CaSO}_4$
- J.** $\text{CO}_3^{2-} + \text{Ca}^{2+} \longrightarrow \text{CaCO}_3$
- K.** $\text{CaCl}_2 + \text{Cu}^{2+} \longrightarrow \text{Ca}^{2+} + \text{CuCl}_2$
- L.** none of the above

Questions On Net Ionic Equations

Joseph A. Mascetta, Mark Kernion



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