

Name _____ Date _____
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REPORT SHEET

LAB

Reaction Rates and Chemical Equilibrium

18

A. Factors That Affect the Rate of a Reaction

A.1 Effect of Temperature

Test Tube Temperature (°C)

1. 51°

2. 92°

Observations (4)

Clear with bubbles

Clear with bubbles

Which test tube cleared first (4)?

Tube 1

Q1 How did an increase in temperature affect the rate of a reaction?

As the temperature increases, particles collide faster increasing the kinetic energy of the molecules, increasing the rate of reaction.

A.2 Effect of Changing Reactant Concentration

Test Tube	Concentration of HCl	Initial time (s)	Final time (s)	Total time (s) = (2) - (1)
1.	1.0 M HCl	0:50	3:42	3:42
2.	2.0 M HCl	0:47	1:47	1:47
3.	3.0 M HCl	0:39	1:39	1:39

Rank the test tubes in order of slowest to fastest.

1 < 2 < 3

Q2 How did an increase in the concentration of HCl cause a difference in the rates of reaction of HCl with Mg?

As the concentration increases, the molecules are increasing and the rate of reaction is enhanced.

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SA Dillow



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