

Colligative Properties of Solutions Worksheet

1. What are colligative properties of solutions? Give examples of three colligative properties.
2. How many particles in solution are produced by each formula unit of potassium carbonate?
3. How many moles of particles would 3 mol Na_2SO_4 give in solution?
4. An equal number of moles of NaCl and CaCl_2 are dissolved in equal volumes of water. Which solution has the lower
 - a. freezing point
 - b. vapor pressure
 - c. boiling point
5. Find the molality of each of the following solutions?
 - a. 2.3 moles of glucose dissolved in 500.0 g of water
 - b. 131 grams of barium nitrate dissolved in 750.0 grams of water
6. Find the boiling points of the following solutions:
 - a. 2.00 m solution of sodium chloride, NaCl
 - b. 1.50 m solution of calcium chloride, CaCl_2
7. Find the freezing points of the following solutions:
 - a. 0.35 moles of sodium chloride dissolved in 900.0 grams of water
 - b. 126.0 grams of strontium nitrate dissolved in 2500.0 grams of water

Properties Of Solutions Guided Answers

Disha Experts

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Disha Experts' text.

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